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## AEE Certified Energy Manager (CEM) Sample Questions (Q31-Q36):

### NEW QUESTION # 31

How much energy (kJ) is required to heat 2 kg of water from 25°C to 2 kg of water at 100°C? (The specific heat of water (C) is 4.2 kJ/kg-°C. Assume no system losses.)

- A. 510 kJ
- **B. 630 kJ**
- C. 420 kJ
- D. 840 kJ

**Answer: B**

### NEW QUESTION # 32

An air conditioning unit cools make-up air from 25°C (dry bulb temperature) and 50% relative humidity down to 13°C. What type of heat has been removed from the air?

- A. No heat energy has been removed
- **B. Both sensible and latent heat**
- C. Sensible heat
- D. Latent heat

**Answer: B**

### NEW QUESTION # 33

A controlled valve is full closed when it gets a signal of 4 mA. The valve is full open when it gets a signal of 20 mA. With a control signal of 8 mA, what is the valve position?

- A. 12.5% open
- B. 45% open
- **C. 25% open**
- D. 75% open
- E. 82% open

**Answer: C**

### NEW QUESTION # 34

Which of the following color temperatures looks more like daylight?

- A. 3,500 K
- B. 2,500 K
- **C. 6,500 K**
- D. 4,500 K

**Answer: C**

### NEW QUESTION # 35

A concentrating solar thermal system collects 5 GJ/h of thermal heat, which is delivered to a 94% efficient hot water generator. The hot water generator is used to raise the water temperature from 15°C to 65°C. The specific heat of water (Cp) is 4.2 kJ/kg-°C. What flow rate (liter/second) of hot water can be delivered from the hot water generator?

- A. 4.2 liter/second
- B. 6.2 liter/second
- C. 12.0 liter/second
- **D. 9.2 liter/second**

**Answer: D**

Explanation:

Step 1: Compute Useful Thermal Energy Output

Given:

\* Solar system delivers 5 GJ/h

\* Hot water generator efficiency = 94%

A screenshot of a math problem AI-generated content may be incorrect.

$$\text{Useful Energy Output} = 5 \times 0.94 = 4.7 \text{ GJ/h}$$

### Step 2: Compute Energy Required to Heat Water

- Specific heat ( $C_p$ ) = 4.2 kJ/kg-°C
- Temperature rise ( $\Delta T$ ) = 65°C - 15°C = 50°C

$$Q = m \times C_p \times \Delta T$$

Since we need flow rate (liters/second), we rearrange for m:

$$m = \frac{Q}{C_p \times \Delta T}$$

$$m = \frac{4.7 \times 10^9}{4.2 \times 50}$$

$$m = \frac{4.7 \times 10^9}{210}$$

$$m = 22,380,952 \text{ g/h} = 22,381 \text{ kg/h}$$

### Step 3: Convert to Liters per Second

- 1 kg of water = 1 liter



$$22,381 \text{ L/h} = \frac{22,381}{3600} = 9.2 \text{ L/s}$$

Thus, the correct answer is C. 9.2 liter/second.

### NEW QUESTION # 36

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